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In the Claims:

1. (Currently Amended) A method of producing chlorine dioxide ( $\text{ClO}_2$ ), comprising:  
reducing alkaline chlorate in a reaction with urea, as a reducing agent, in a mineral acid medium  
to produce  $\text{ClO}_2$  at or below normal atmospheric pressure.
2. (Original) A method according to claim 1 wherein the alkaline chlorate is  $\text{NaClO}_3$  or  $\text{KClO}_3$ .
3. (Original) A method according to claim 2 wherein the mineral acid is selected from the group  
consisting essentially of sulfuric acid, nitric acid, phosphoric acid and hydrochloric acid.
4. (Cancel)
5. (Currently Amended) A method according to claim ~~[[4]]~~ 3 wherein below atmospheric  
pressure is about 100 to about 400 mm Hg.
6. (Original) A method according to claim 5 comprising controlling a mol ratio of alkaline  
chlorate and mineral acid to be between about 0.2 to about 0.8 in the reaction.
7. (Original) A method according to claim 6 comprising controlling a mineral acid concentration  
of the reaction to be between about 3  $\text{mol-L}^{-1}$  to about 10  $\text{mol-L}^{-1}$ .
8. (Original) A method according to claim 7 comprising controlling a mol ratio of alkaline  
chlorate and Urea in the reaction to be between about 2 to about 4, except that in hydrochloric  
acid medium, the mol ratio is about 0.84.
9. (Original) A method according to claim 8 comprising controlling a temperature of the reaction  
to be between about 40°C to about 90°C.
10. (Original) A method according to claim 9 comprising controlling a temperature of the  
reaction to be between about 55°C to about 85°C.

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11. (Original) A method of producing chlorine dioxide ( $\text{ClO}_2$ ) and at least one inorganic salt, comprising: reducing alkaline chlorate in a reaction with urea as reducing agent in a mineral acid medium to produce  $\text{ClO}_2$ ; then, adding an alcohol as salting out reagent to the medium; and, collecting an inorganic salt.

12. (Original) A method according to claim 11 wherein the alcohol is selected from the group consisting essentially of methanol, ethanol, and propanol.

13. (Original) A method according to claim 12 wherein the alkaline chlorate is  $\text{NaClO}_3$  or  $\text{KClO}_3$ , and the mineral acid is selected from the group consisting essentially of sulfuric acid, nitric acid, phosphoric acid and hydrochloric acid.

14. (Original) A method according to claim 13 wherein the inorganic salt is selected from the group consisting essentially of potassium sulfate, potassium nitrate, and sodium nitrate.

15. (Currently Amended) A method of producing chlorine dioxide ( $\text{ClO}_2$ ) and at least one binary/ternary compound fertilizer, comprising: reducing alkaline chlorate in a reaction with urea as a reducing agent in a mineral acid medium to produce  $\text{ClO}_2$ ; then, adding urea  $[(\text{I})]$  or other nitrogenous fertilizer  $[(\text{J})]$  selected from the group consisting essentially of phosphorus ore powder  $[(\text{I})]$  and phosphate fertilizer  $[(\text{amount based on the acid concentration of mother solution})]$ , and/or kali salt to the medium; and, collecting a compound fertilizer.

16. (Original) A method according to claim 15 wherein the alkaline chlorate is  $\text{NaClO}_3$  or  $\text{KClO}_3$  and the mineral acid is selected from the group consisting essentially of sulfuric acid, nitric acid, phosphoric acid and hydrochloric acid.

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